**SODIUM CHLORIDE TESTING**

**Aim:**  To test different strengths of sodium chloride solutions, as to see how they would effect salinity levels in water.

**Hypothesis:** The added sodium chloride solution will have a directly proportional effect on salinity levels.

**Materials:**

* Potassium chromate indicator
* Silver nitrate standard solution (0.1 M for sea water samples)
* Distilled water in wash bottle
* Burette
* Pipette
* Pipette filler
* 250 mL conical flask
* Small funnel
* 0.01M for fresh water samples
* Sodium Chloride Solution (1%, 2%, 4% and 8%)

**Procedure:**

1. The burette was rinsed with silver nitrate solution, then fixed in the clamp and stand
2. The burette was filled to the 0mL level
3. We then rinsed the flask with distilled water then using the pipette, we transferred 5mLs of the water sample to the flask
4. 4 drops of the potassium chromate indicator solution was added and swirled to mix
5. Then we titrate until a red-brown end-point was reached
6. The amount of titre was recorded (volume of AgNO3 used) and the above steps were repeated twice and the average was found
7. The moles/L of sodium chloride was calculated, then converted to mg/L then to ppm
8. Steps 3-7 were then repeated twice for each different sodium chloride solution concentration and an average was calculated